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nature of matter dynamic equilibrium electron affinity electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements empirical formula energy changes and intermolecular attractions energy of revolving electron experimental techniques filter paper filtration crucibles fundamental particles gas laws gas properties graham s law grahams law of diffusion heisenberg s uncertainty principle hexagonal close packing higher ionization energies hydrogen bonding hydrogen spectrum ideal gas constant ideal gas density ideality deviations intermolecular forces ionic radius ionization energies ionization energy isotopes kinetic interpretation of temperature kinetic molecular theory of gases lewis concept liquefaction of gases liquid crystals liquids properties london dispersion forces magnetic quantum number mass of electron mass spectrometer metallic crystals properties metallic solids metals structure molar volume molecular ions molecular solids molecules moles moseley law neutron properties non ideal behavior of gases orbital concept partial pressure calculations phase changes energies photons wave number planck s quantum theory plasma state positive and negative ions pressure units properties of cathode rays covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der waals equation vapor pressure and spectrum

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2011-06

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