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explain the difference between a nonpolar covalent bond a polar covalent bond and an ionic bond from its position in the periodic table determine which atom in each pair is more electronegative br or cl

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in ionic bonding the two ways atoms can get the same configuration as a noble gas is by or electrons

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there are two major types of chemical bonds ionic bonds and covalent bonds an ionic bond is a bond that results from the electrostatic attraction force between ions of opposite charges ionic bonds apply to ionic compounds such as sodium chloride nacl

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exactly what is a chemical bond and what observable properties can we use to distinguish one kind of bond from another this is the first of ten lessons that will help familiarize you with the fundamental concepts of this very broad subject

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are gained

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provided 1 a a chemical bond between atoms results from the attraction between the
valence electrons and of different atoms a nuclei c isotopes b inner electrons d lewis
structures 2 b a covalent bond consists of a a shared electron c two

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question 2 state the terms defined in each case a bond formed by a a shared pair of electrons each bonding atom contributing one electron to the pair b a shared pair of electrons with both electrons coming from the same atom answer a covalent bond b coordinate bond 2009 question 1

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the lewis electron system is a simplified approach for understanding bonding in covalent and ionic compounds why do chemists still find it useful is a lewis dot symbol an exact representation of the valence electrons in an atom or ion explain your answer

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