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an aqueous solution is a chemical solution in which water is the solvent learn how to identify write and compare aqueous solutions and solve problems involving their colligative properties an aqueous solution is a solution in which the solvent is water it is mostly shown in chemical equations by appending aq to the relevant chemical formula for example a solution of table salt also known as sodium chloride nacl in water would be represented as na aq cl aq an aqueous solution is any solution in which water h₂o is the solvent in a chemical equation the symbol aq follows a species name to indicate that it is in aqueous solution for example dissolving salt in water has the chemical reaction an aqueous solution is water that contains one or more dissolved substance the dissolved substances in an aqueous solution may be solids gases or other liquids in order to be a true solution a mixture must be stable in a solution a solute the substance present in the lesser amount is dispersed in a solvent the substance present in the greater amount aqueous solutions contain water as the solvent whereas nonaqueous solutions have solvents other than water aqueous solutions are all around us and even inside of us aqueous solutions are homogeneous mixtures that contain water as their main component learn about solutes solvents electrolytes and why water is known as the universal solvent in a solution a solute the substance present in the lesser amount is dispersed in a solvent the substance present in the greater amount aqueous solutions contain water as the solvent whereas nonaqueous solutions have solvents other than water the water in oceans and rivers counts as aqueous solution carrying a broad diversity of salts and organic compounds to support the organisms that inhabit them aqueous solutions with a soap or detergent solute are used to remove dirt and germs from dishes clothes surfaces and human skin in a solution a solute the substance present in the lesser amount is dispersed in a solvent the substance present in the greater amount aqueous solutions contain water as the solvent whereas nonaqueous solutions have solvents other than water there are three important classes of reactions which occur in aqueous solution precipitation reactions acid base reactions and redox reactions topic hierarchy in aqueous solutions water is the solvent even if it is in lesser quantities and in this class we will be dealing mostly with aqueous solutions the solute can be of a different phase than the solvent define what makes an aqueous solution explain the behavior of ionic compounds in water calculate molarity of a solution predict the solubility of different combinations of solutes and solvents given the chemical properties of those compounds this chapter introduces techniques for preparing and analyzing aqueous solutions for balancing equations that describe reactions in solution and for solving problems using solution stoichiometry by the time you complete this chapter you will know enough about aqueous solutions to explain what causes acid rain why acid rain is harmful and a substance will form an aqueous solution or not it depends on the nature of its chemical bonds aqueous is a term used to define a system that involves water the word aqueous is also applicable to describe a solution or mixture in which water is the solvent aqueous solutions not only cover most of earth's surface as seawater but they also play vital roles in both environmental and industrial processes and are essential to the proper functioning an aqueous solution is a solution in which the solvent is water whereas in

a nonaqueous solution the solvent is a substance other than water familiar examples of nonaqueous solvents are ethyl acetate used in nail polish removers and turpentine used to clean paint brushes when water is the solvent for a reaction the reaction is said to occur in aqueous solution which is denoted by the abbreviation aq following the name of a chemical species in a reaction three important types of reactions in water are precipitation acid base and oxidation reduction reactions precipitation reactions key points an arrhenius acid is any species that increases the concentration of h in aqueous solution an arrhenius base is any species that increases the concentration of oh in aqueous solution in aqueous solution h ions immediately react with water molecules to form hydronium ions H_3O^+ aqueous is a term used to describe a system which involves water the word aqueous is also applied to describe a solution or mixture in which water is the solvent when a chemical species has been dissolved in water this is denoted by writing aq after the chemical name chemistry 101 aqueous solutions class lectures aqueous solutions are a fundamental concept in chemistry that involves the dissolution of a solute in a solvent to form a homogeneous mixture the solute can be a solid liquid or gas while the solvent is typically water

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