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problem 1 calculate the decrease in temperature in celsius when 2 00 l at 21 0 c is compressed to 1 00 l solution 2 00 l 294 0 k 1 00 l x cross multiply to get 2x 293 x 147 0 k converting 147 0 k to celsius we find 126 0 c for a total decrease of 147 0 c from 21 0 c to 126 0 c

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charles s law states that increasing the temperature of a gas at constant pressure increases its volume charles s law formula there are a few ways to state charles law as a formula v t v t k v kt v 1 t 1 v 2 t 2 v 2 v 1 t 2 t 1 v 1 t 2 v 2 t 1 here t is absolute temperature v is volume and k is a non zero constant

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charles s law of gases indicates that at a constant pressure the volume of a gas is proportional to the temperature this is calculated by the following equation our first step to solving this equation will be to convert the given temperature to kelvin using this temperature and the given volumes we can solve for the final temperature of

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the charle s law shows the correlation between the temperature and the volume of an ideal gas let s discuss it by using a pump with a freely moving plunger that is filled with some gas in the experiment we are going to heat up the pump and monitor how the pressure and volume are affected

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