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the rate or speed at which a reaction occurs depends on the frequency of successful collisions remember a successful collision occurs when two reactants collide with enough energy and with the right orientation

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explore how factors like temperature and concentration shift reaction rates and witness the balancing act of reversible reactions use le châtelier s principle to predict how changes impact chemical equilibrium and discover how catalysts are a vital part of every living system

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at equilibrium the rates of the forward and reverse processes are identical but the rate constants are generally different to see how this works consider the simplified reaction $a + b \rightleftharpoons c + d$ in the following three scenarios

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equilibrium is when the rate of the forward reaction equals the rate of the reverse reaction all reactant and product concentrations are constant at equilibrium given a reaction $a + b \rightleftharpoons c + d$ the equilibrium constant K_c also called K or K_{eq} is defined as follows $K_c = \frac{[c][d]}{[a][b]}$

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however certain changes in the reaction conditions can have an effect on the rate of a given chemical reaction collision theory can be utilized to explain these rate effects 7 7 reversible reactions and chemical equilibrium 7 8 equilibrium equations and equilibrium constants

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introduction to equilibrium learn dynamic equilibrium direction of reversible reactions types of equilibrium constants learn writing equilibrium constant and reaction quotient expressions the equilibrium constant k calculating equilibrium constant k_p using partial pressures worked examples calculating equilibrium constants practice

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the rate of a chemical reaction is the change in the amount of a reactant or product of the reaction with time the rates of chemical reactions differ enormously depending on the reaction and conditions rates of chemical change must be measured for each chemical reaction under given conditions

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when the rates of the forward and reverse reactions are equal the concentrations of the reactant and product species remain constant over time and the system is at equilibrium the relative concentrations of reactants and products in equilibrium systems vary greatly some systems contain mostly products at equilibrium some contain mostly

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chemical equilibrium condition in the course of a reversible chemical reaction in which no net change in the amounts of reactants and

products occurs a reversible chemical reaction is one in which the products as soon as they are formed react to produce the original reactants

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eventually the rates of the forwards and reverse reactions become equal and equilibrium is reached particles are still reacting but as a and b react to produce c and d c and d react to produce a and b maintaining a constant concentration

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explore what makes a reaction happen by colliding atoms and molecules design experiments with different reactions concentrations and temperatures when are reactions reversible what affects the rate of a reaction

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reaction rates and equilibrium 593 understanding by design students are building toward understanding how matter and energy behave in chemical reactions by applying the concepts of reaction rates and equilibrium performance goals at the end of chapter 18 students will be able to answer the essential questions by applying their knowledge of

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equilibrium we have already seen that a reversible reaction is one that occurs in both directions when during the course of reaction the rate of the forward reaction equals the rate of the reverse reaction then the overall reaction is said to be in a state of equilibrium

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the description of how a system responds to a stress to equilibrium has become known as le châtelier s principle when a chemical system that is at equilibrium is disturbed by a stress the

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law of mass action equilibrium constant K_{eq} the law of mass action for a particular reaction has a specific numerical value for a reaction at a

given temperature the reaction must contain molecules that are all in the same phase gas or liquid

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this chapter examines the idea of reversible reactions and their equilibrium positions significant emphasis is placed on how equilibrium and the rate of a reaction can be affected by altering temperature pressure and concentrations

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