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under the arrhenius definition acids and bases come together to form water neutralizing each other in the process $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O}$ l bronsted lowry definition of acids and bases acid donates a H^+ ion base accepts a H^+ ion

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to identify the strongest base we can determine their weakest conjugate acid the conjugate acids of CH_3NH_2 and S^{2-} are CH_4 NH_3 and HS^- respectively next we consider that acidity increases with positive charge on the molecule thus ruling out that S^{2-} is the weakest base

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pancake 2 days ago great question first of all NaCl is not an acid instead it is considered a salt salts contain a cation from a base and an anion from an acid NaOH a base reacts with HCl an acid to produce NaCl and H_2O what makes a substance dangerous is its hydrogen ion concentration and strength

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1 identifying acids and bases write the chemical equation for the ionization reaction and the corresponding K_a equilibrium expression for each acid a HCN b H_2S c $\text{C}_5\text{H}_5\text{NH}_2$ d HF a answer this content is available to registered users only click here to register

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quiz 1 acids bases and pH acid base equilibria learn weak acid base equilibria conjugate acid base pairs relationship between K_a and K_b K_a and acid strength weak acid equilibrium weak base equilibrium acid base properties of salts pH of salt solutions this unit is part of the chemistry library

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1 use table k and table l to help you identify the rules for determining whether a substance is an acid a base or a salt based on the formula underline all the acids circle bases and box in salts purple leave the covalent substances alone NH_3 NaCl CH_3OH H_2SO_4 Ca(OH)_2 CH_4 NH_4Br HCl Na_2SO_4 HNO_3 CH_3

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hypochlorous acid dissociates in water to create hydronium ions and hypochlorite ions $\text{HOClo}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OCl}^-$ suppose that additional hypochlorite ions are added to the

solution which of the following correctly describes the resultant effect on the concentration of H^+ choose 1 answer it depends on the number of hydronium ions

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by anne marie helmenstine ph d updated on may 05 2019 1 which of the following statements is true concerning acids and bases acids and bases don't react with each other acids mixed with bases neutralize each other acids mixed with bases make stronger bases acids mixed with bases make stronger acids

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answers arrhenius acid a compound that increases the concentration of hydrogen ion H^+ in aqueous solution arrhenius base a compound that increases the concentration of hydroxide ion OH^- in aqueous solution the reaction of an acid and a base

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the concept of conjugate acids and bases is best understood by considering what happens when a substance behaves as a brønsted lowry acid or a brønsted lowry base by reversing the reaction in which a substance acts as a proton donor we see that the product is itself a proton acceptor

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acids and bases 1 worksheet name section student id work in groups on these problems you should try to answer the questions without referring to your textbook

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acids and bases worksheet 1 answers 1 classify each of the following substances as an acid or a base according to the arrhenius definition a acid b base c base d acid e base 2 classify each of the following substances as an acid a base or amphoteric substance according to the brønsted lowry definition a acid b acid c acid d

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a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste an acid is basically a molecule which can donate an H^+ ion and can remain energetically favourable after a loss of H^+ acids are known to turn blue litmus red

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many acids and bases are weak that is they do not ionize fully in aqueous solution a solution of a weak acid in water is a mixture of the nonionized acid hydronium ion and the conjugate base of the acid with the nonionized acid present in the greatest concentration

14 acids and bases Jan 16 2023

choose from Cu , Ag , Au , Pt , Pd or Hg 6 write balanced chemical equations for the following reactions of acids and bases a aluminum metal with dilute nitric acid $2\text{Al} + 6\text{HNO}_3 \text{ aq} \rightarrow 2\text{Al}(\text{NO}_3)_3 \text{ aq} + 3\text{H}_2 \text{ g}$ b calcium hydroxide solution with acetic acid $\text{Ca}(\text{OH})_2 \text{ aq} + 2\text{CH}_3\text{COOH} \text{ aq} \rightarrow \text{Ca}(\text{CH}_3\text{COO})_2 \text{ aq}$

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syllabus checklist names and formulae of common acid and bases brønsted lowry theory of acid and bases a brønsted lowry acid is a proton H^+ donor a brønsted lowry base is a proton H^+ acceptor a proton is a H^+ ion but exists in solution as H_3O^+ aq

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