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 \mathbb{Z} \mathbb{Z} for a first order reaction the units of k are 1 s for a second order reaction the units of k are 1 mol s for an elementary reaction the rate law order is identical to the stoichiometric coefficients for heterogeneous reactions partial pressures are

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 $\ensuremath{\mathbb{Z}}$ $\ensuremath{\mathbb{Z}}$ here are the main ones reactor design is confined to simple reactions most textbooks focus on the design of chemical reactors with single reactions only a brief discussion is devoted to reactors with multiple reactions in practice of course engineers rarely encounter chemical reactors with single reactions

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