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a dilute solution is one in which there is a relatively small amount of solute dissolved in the solution a concentrated solution contains a relatively large amount of solute these two terms do not provide any quantitative information actual numbers but they are often useful in comparing solutions in a more general sense a common method of making a solution of a given concentration involves taking a more concentrated solution and adding water until the desired concentration is reached this process is known as dilution decide on the concentration of the obtained solution let's say you want it to be equal to 20 mm 20 mm input all this data into the dilution equation $c_1 v_1 = c_2 v_2$ $10 \text{ mm} \cdot 10 \text{ ml} = 20 \text{ mm} \cdot v_2$ $v_2 = 5 \text{ ml}$ for diluting solutions in lab experiments the formal formula for calculating a dilution is $c_1 v_1 = c_2 v_2$ where c_1 and c_2 represent the concentrations of the initial and final solutions respectively and v_1 and v_2 represent their volumes method 1 accurately diluting concentrates via dilution equation download article 1 to dilute a solution means to add more solvent without the addition of more solute of course the resulting solution is thoroughly mixed so as to ensure that all parts of the solution are identical the fact that the solute amount stays constant allows us to develop calculation techniques first we write dilution equation c_1 is the concentration of the stock solution v_1 is the volume to be removed i.e. aliquoted from the concentrated stock solution c_2 is the final concentration of the diluted solution v_2 is the final volume of the diluted solution what are dilutions dilution is an important experimental technique for creating solutions of a desired concentration a solution is a mixture of components that is homogenous at the molecular level the solution components are two or more pure substances that were mixed to form the solution how to calculate units of concentration once you have identified the solute and solvent in a solution you are ready to determine its concentration concentration may be expressed several different ways using percent composition by mass volume percent mole fraction molarity molality or normality how to calculate concentration of solution when it's diluted what are some examples of dilution calculations what would be the concentration of a solution made by adding 250 ml of water to 450 ml of 4.2 M KOH what would be the concentration of a solution made by diluting 450 ml of 4.2 M KOH to 250 ml diluting a solvent means to lower the concentration by adding more of the solvent to the solution this can be done by adding solvent to the main bulk of the solution or taking some volume of the

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solution and adding solvent to it the molarity or molar concentration of a solute is defined as the number of moles of solute per liter of solution not per liter of solvent what is a mole this video on the mole and avogadro s number molarity mol solute l of solution why is the volume of the solution different from the volume of the solvent dilution is the process of decreasing the concentration of a solute in a solution usually simply by mixing with more solvent like adding more water to the solution to dilute a solution means to add more solvent without the addition of more solute chemistry solutions dilution definition what is dilution dilution is the process of lowering the concentration of a solute in a solution by simply adding more solvent to the solution such as water diluting a solution entails adding more solvent without adding more solute 1 answer sorted by 2 the answer is basically correct see note at the end of my answer about significant figures but there is a simpler method using a dilution factor the original volume of your solution was 2 ml and the final volume was 12 ml so the dilution factor is simply $2/12$ the dilution ratio calculator tells you how much solute and solvent you need to get the desired dilution ratio our tool has a built in volume conversion so you will be able to perform your calculations using any units you want this molarity calculator is a tool for converting the mass concentration of any solution to molar concentration or recalculating grams per ml to moles you can also calculate the mass of a substance needed to achieve a desired molarity this article will provide you with the molarity definition and the molarity formula

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a common method of making a solution of a given concentration involves taking a more concentrated solution and adding water until the desired concentration is reached this process is known as dilution

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decide on the concentration of the obtained solution let s say you want it to be equal to 20 m m 20 mm 20 mm input all this data into the dilution equation $c_1 v_1 = c_2 v_2$ $1 \cdot 20 = 10 \cdot 3 \cdot 0.5$ $v_1 = 1.5$ 0.01 10 ml end split 1 v 1 v 1 20 10 3 0.5 0.01 10 ml

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for diluting solutions in lab experiments the formal formula for calculating a dilution is $c_1 v_1 = c_2 v_2$ where c_1 and c_2 represent the concentrations of the initial and final solutions respectively and v_1 and v_2 represent their volumes method 1 accurately diluting concentrates via dilution equation download article 1

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to dilute a solution means to add more solvent without the addition of more solute of course the resulting solution is thoroughly mixed so as to ensure that all parts of the solution are identical the fact that the solute amount stays constant allows us to develop calculation techniques first we write

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dilution equation c_1 is the concentration of the stock solution v_1 is the volume to be removed i e aliquoted from the concentrated stock solution c_2 is the final concentration of the diluted solution v_2 is the final volume of the diluted solution

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what are dilutions dilution is an important experimental technique for creating solutions of a desired concentration a solution is a mixture of components that is homogenous at the molecular level the solution components are two or more pure substances that were mixed to form the solution

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how to calculate units of concentration once you have identified the solute and solvent in a solution you are ready to determine its concentration concentration may be expressed several different ways using percent composition by mass volume percent mole fraction molarity molality or normality

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diluting a solvent means to lower the concentration by adding more of the solvent to the solution this can be done by adding solvent to the
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main bulk of the solution or taking some volume of the solution and adding solvent to it

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the molarity or molar concentration of a solute is defined as the number of moles of solute per liter of solution not per liter of solvent what is a mole this video on the mole and avogadro s number molarity mol solute l of solution why is the volume of the solution different from the volume of the solvent

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dilution is the process of decreasing the concentration of a solute in a solution usually simply by mixing with more solvent like adding more water to the solution to dilute a solution means to add more solvent without the addition of more solute

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chemistry solutions dilution definition what is dilution dilution is the process of lowering the concentration of a solute in a solution by simply adding more solvent to the solution such as water diluting a solution entails adding more solvent without adding more solute

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1 answer sorted by 2 the answer is basically correct see note at the end of my answer about significant figures but there is a simpler method using a dilution factor the original volume of your solution was 2 ml and the final volume was 12 ml so the dilution factor is simply $212 \ 2 \ 12$

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the dilution ratio calculator tells you how much solute and solvent you need to get the desired dilution ratio our tool has a built in volume conversion so you will be able to perform your calculations using any units you want

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this molarity calculator is a tool for converting the mass concentration of any solution to molar concentration or recalculating grams per ml to moles you can also calculate the mass of a substance needed to achieve a desired molarity this article will provide you with the molarity definition and the molarity formula

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