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## **3 1 chemical equations chemistry libretexts**

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a chemical reaction is described by a chemical equation an expression that gives the identities and quantities of the substances involved in a reaction a chemical equation shows the starting compound s the reactants on the left and the final compound s the products on the right separated by an arrow

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derive chemical equations from narrative descriptions of chemical reactions write and balance chemical equations in molecular total ionic and net ionic formats

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a balanced chemical equation shows the same number of each type of atom on both sides of the arrow

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to balance a chemical equation enter an equation of a chemical reaction and press the balance button the balanced equation will appear above use uppercase for the first character in the element and lowercase for the second character examples fe au co br c o n f ionic charges are not yet supported and will be ignored

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balance the following chemical equation  $\text{mg}(\text{oh})_2 + \text{hcl} \rightarrow \text{mgcl}_2 + \text{h}_2\text{o}$  note all reactants and products require a coefficient of at least one learn for free about math art computer programming economics physics chemistry biology medicine finance history and more

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example 1 how many oxygen atoms are represented in the formula  $\text{agn}_3\text{o}_3$  oxygen atoms explain please o example 2 how many oxygen atoms are represented in the formula  $\text{mg}(\text{no}_3)_2$

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reactants and products to describe a chemical reaction we need to indicate what substances are present at the beginning and what substances are present at the end the substances that are present at the beginning are called reactants and the substances present at the end are called products

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balance chemical equations with fun and interactive simulations learn the rules test your knowledge and challenge yourself with different levels

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learn how to balance chemical equations by using the law of conservation of mass and the coefficients of reactants and products practice with different levels of difficulty and get immediate feedback compare your results with real life examples and simulations of chemical reactions

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write and balance the chemical equation described by exercise 4 the formula for propane is  $C_3H_8$  balance  $NaClO_3 \rightarrow NaCl + O_2$  balance  $N_2H_2 + N_2H_4 \rightarrow Al_2O_3$  balance  $C_2H_4 + O_2 \rightarrow CO_2 + H_2O$  how would you write the balanced chemical equation in exercise 10 if

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which reactant is missing from the equation shown three hydrogen atoms  $2Na_2S_2O_3$  what is the product of the equation shown  $Na_2S_2O_3$  study with quizlet and memorize flashcards containing terms like which of the following is a chemical reaction  $2H_2 + O_2 \rightarrow 2H_2O$  in the equation shown what are the reactants sulfuric

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the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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how do you know if a chemical equation is balanced what can you change to balance an equation play a game to test your ideas

## **what are chemical equations detailed explanation examples**

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table of content what are chemical equations ionic chemical equations faqs on chemical equations what is a chemical equation chemical equations make use of symbols to represent factors such as the direction of the reaction and the physical states of the reacting entities

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mg alcl<sub>3</sub> al mgcl<sub>2</sub> b 3mg 2alcl<sub>3</sub> 2al 3mgcl<sub>2</sub> which set of coefficients would balance the following chemical equation h<sub>3</sub>po<sub>4</sub> rboh h<sub>2</sub>o rb<sub>3</sub>po<sub>4</sub> d h<sub>3</sub>po<sub>4</sub> 3rboh 3h<sub>2</sub>o rb<sub>3</sub>po<sub>4</sub> silver ag has a molar mass of 107 8682g nitrogen has a molar mass of 14 0067g and oxygen has a molar mass of 15 9994g

## 5 1 1 practice problems writing and balancing chemical

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answer c ce h<sub>2</sub> g ce i<sub>2</sub> s rightarrow ce 2hi s answer d ce 4fe s ce 3o<sub>2</sub> g rightarrow ce 2fe<sub>2</sub>o<sub>3</sub> s answer e ce 2na s ce 2h<sub>2</sub>o l rightarrow ce 2naoh aq ce h<sub>2</sub> g answer f ce nh<sub>4</sub> 2cr<sub>5</sub>o<sub>7</sub> s rightarrow ce cr<sub>2</sub>o<sub>3</sub> s ce n<sub>2</sub> g ce 4h<sub>2</sub>o g answer g ce p<sub>4</sub> s ce 6cl<sub>2</sub> g rightarrow ce 4pcl<sub>3</sub> l answer h

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examples of complete chemical equations to balance fe cl 2 fecl 3 kmno 4 hcl kcl mncl 2 h 2 o cl 2 k 4 fe cn 6 h 2 so 4 h 2 o k 2 so 4 feso 4 nh 4 2 so 4 co c 6 h 5 cooh o 2 co 2 h 2 o k 4 fe cn 6 kmno 4 h 2 so 4 khso 4 fe 2 so 4 3 mnso 4 hno 3 co 2 h 2 o

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question 2 write the balanced equation for the following chemical reactions i hydrogen chlorine hydrogen chloride ii barium chloride aluminium sulphate barium sulphate aluminium chloride iii sodium water sodium hydroxide hydrogen answer i h 2 cl 2 2hcl

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