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what would the frequency of this light be  $470\text{ nm} = 4.70 \times 10^{-7}\text{ m}$   $v = c/\lambda = 3 \times 10^8\text{ m/s} / 4.70 \times 10^{-7}\text{ m} = 6.38 \times 10^{14}\text{ s}^{-1} = 6.38 \times 10^{14}\text{ Hz}$   $c$  how much energy would a single photon of this light possess express your answer in joules  $e = h\nu = 6.626 \times 10^{-34}\text{ J}\cdot\text{s} \times 6.38 \times 10^{14}\text{ s}^{-1} = 4.23 \times 10^{-19}\text{ J}$  d how much energy would a mole of such photons possess

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