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chapter 15 3 solving equilibrium problems chemistry libretxts Mar 22 2024 at equilibrium the mixture contained 0.00272 mol/L NH<sub>3</sub> what is K for the reaction N<sub>2</sub> + 3H<sub>2</sub> ⇌ 2NH<sub>3</sub> at this temperature what is K p answer K = 0.105 K<sub>p</sub> = 2.61 × 10<sup>5</sup>

**2 e chemical equilibrium practice problems with answers** Feb 21 2024 answers the system is not at equilibrium at each of these higher pressures to reach equilibrium the reaction will proceed to the right to decrease the pressure because the equilibrium partial pressure is less than the total pressure K<sub>p</sub> =  $\frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3} = \frac{15.20^2}{2.19 \times 17.65^3} = 13.34 \times 10^{-5}$

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**ncert solutions for class 11 chemistry chapter 7 equilibrium** Nov 06 2022 q 17 k<sub>p</sub> = 0.04 atm at 899 K for the equilibrium shown below what is the equilibrium concentration of C<sub>2</sub>H<sub>6</sub> when it is placed in a flask at 4.0 atm pressure and allowed to come to equilibrium C<sub>2</sub>H<sub>6</sub> + g C<sub>2</sub>H<sub>4</sub>

g h 2 g ans let p be the pressure exerted by ethene and hydrogen gas each at equilibrium now according to the reaction  
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