

Download free Acidic solution definition (Read Only)

an acidic solution is any aqueous solution which has a $pH < 7$ an aqueous solution is one which has water as its solvent while it is never a good idea to taste an unknown solution acidic solutions are sour in contrast to alkaline solutions which are soapy learn how acids and bases affect the concentration of hydrogen ions in water and how pH measures acidity or basicity find out how buffers help maintain stable pH in biological systems learn how to calculate and interpret pH and pOH values for aqueous solutions pH is a measure of acidity and pOH is a measure of basicity based on the concentrations of H^+ and OH^- ions learn the definition of Arrhenius acids and bases and how they react in aqueous solution an Arrhenius acid is any species that increases the concentration of H^+ ions and an Arrhenius base is any species that increases the concentration of OH^- ions according to Brønsted and Lowry an acid a substance with at least one hydrogen atom that can dissociate to form an anion and an H^+ ion a proton in aqueous solution thereby forming an acidic solution is any substance that can donate a proton and a base a substance that produces one or more hydroxide ions OH^- and a cation when an acidic solution is a liquid mixture that occurs when hydrogen ions are released when combined with water this definition is called the Brønsted Lowry theory acids acid any substance that in water solution tastes sour changes the colour of certain indicators e.g. reddens blue litmus paper reacts with some metals e.g. iron to liberate hydrogen reacts with bases to form salts and promotes certain chemical reactions acid catalysis this is known as the pH scale and is the range of values from 0 to 14 that describes the acidity or basicity of a solution you can use pH to quickly determine whether a given aqueous solution is acidic basic or neutral learn about solutions solubility molarity and pH in this high school chemistry unit an acidic solution is a solution that contains hydrogen ions H^+ and has a pH less than 7 aqueous solutions at 25 °C with a pH less than 7 are acidic while those with a pH greater than 7 are basic or alkaline a pH level of 7.0 at 25 °C is defined as neutral because the concentration of H^+ equals the concentration of OH^- in pure water an acid is a chemical species that donates protons or hydrogen ions and or accepts electrons most acids contain a hydrogen atom bonded that can release dissociate to yield a cation and an anion in water the higher the concentration of hydrogen ions produced by an acid the higher its acidity and the lower the pH of the solution acid base reaction a type of chemical process in which one or more hydrogen ions are exchanged between species that may be neutral molecules such as water or acetic acid or electrically charged ions such as ammonium carbonate or hydroxide a solution with a pH less than 7 is considered acidic a solution with a pH greater than 7 is considered basic or alkaline the Arrhenius definition states that an acid produces H^+ in solution and a base produces OH^- this theory was developed by Svante Arrhenius in 1883 later two more sophisticated and general theories were proposed these are the Brønsted Lowry and the Lewis definitions of acids and bases the Lewis theory is discussed elsewhere acid base properties of salts pH of salt solutions this unit is part of the chemistry library browse videos articles and exercises by topic solutions having a value of pH ranging from 0 to 7 on the pH scale are termed as acidic and the value of pH ranging from 7 to 14 on pH scale are known as basic solutions solutions having the value of pH equal to 7 on pH scale are known as neutral solutions an acidic solution and a basic solution react together in a neutralization reaction that also forms a salt acid base reactions require both an acid and a base in Brønsted Lowry an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste the earliest

definition of acids and bases is arrhenius s definition which states that an acid is a substance that forms hydrogen ions H^+ when dissolved in water and a base is a substance that forms hydroxide ions OH^- when dissolved in water the ph scale is used to rank solutions in terms of how acidic or how basic they are it indicates the concentration of hydrogen ions H^+ and hydroxide ions OH^- in a solution these ion concentrations are equal in pure water which has a ph of 7

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an acidic solution is any aqueous solution which has a $\text{pH} < 7$ an aqueous solution is one which has water as its solvent while it is never a good idea to taste an unknown solution acidic solutions are sour in contrast to alkaline solutions which are soapy

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according to brønsted and lowry an acid a substance with at least one hydrogen atom that can dissociate to form an anion and an H^+ ion a proton in aqueous solution thereby forming an acidic solution is any substance that can donate a proton and a base a substance that produces one or more hydroxide ions OH^- and a cation when

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an acidic solution is a liquid mixture that occurs when hydrogen ions are released when combined with water this definition is called the brønsted lowry theory acids

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acid any substance that in water solution tastes sour changes the colour of certain indicators e g reddens blue litmus paper reacts with some metals e g iron to liberate hydrogen reacts with bases to form salts and promotes certain chemical reactions acid catalysis

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this is known as the ph scale and is the range of values from 0 to 14 that describes the acidity or basicity of a solution you can use ph to quickly determine whether a given aqueous solution is acidic basic or neutral

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learn about solutions solubility molarity and ph in this high school chemistry unit an acidic solution is a solution that contains hydrogen ions h and has a ph less than 7

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aqueous solutions at 25 c with a ph less than 7 are acidic while those with a ph greater than 7 are basic or alkaline a ph level of 7 0 at 25 c is defined as neutral because the concentration of h³ o equals the concentration of oh in pure water

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an acid is a chemical species that donates protons or hydrogen ions and or accepts electrons most acids contain a hydrogen atom bonded that can release dissociate to yield a cation and an anion in water the higher the concentration of hydrogen ions produced by an acid the higher its acidity and the lower the ph of the solution

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acid base reaction a type of chemical process in which one or more hydrogen ions are exchanged between species that may be neutral molecules such as water or acetic acid or electrically charged ions such as ammonium carbonate or hydroxide

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a solution with a ph less than 7 is considered acidic a solution with a ph greater than 7 is considered basic or alkaline

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the arrhenius definition states that an acid produces h in solution and a base produces oh this theory was developed by svante arrhenius in 1883 later two more sophisticated and general theories were proposed these are the brønsted lowry and the lewis definitions of acids and bases the lewis theory is discussed elsewhere

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solutions having a value of ph ranging from 0 to 7 on the ph scale are termed as acidic and the value of ph ranging from 7 to 14 on ph scale are known as basic solutions solutions having the value of ph equal to 7 on ph scale are known as neutral solutions

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an acidic solution and a basic solution react together in a neutralization reaction that also forms a salt acid base reactions require both an acid and a base in brønsted lowry

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an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste

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the earliest definition of acids and bases is arrhenius s definition which states that an acid is a substance that forms hydrogen ions h when dissolved in water and a base is a substance that forms hydroxide ions oh when dissolved in water

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the ph scale is used to rank solutions in terms of how acidic or how basic they are it indicates the concentration of hydrogen ions h and hydroxide ions oh in a solution these ion concentrations are equal in pure water which has a ph of 7

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