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chemical equilibrium when a chemical reaction takes place in a container which prevents the entry or escape of any of the substances involved in the reaction the quantities of these components change as some are consumed and others are formed

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this unit explores the how and why of chemical equilibrium learn about reversible reactions the equilibrium constant le châtelier s principle solubility equilibria and more practice what you ve learned and study for the ap chemistry exam with more than 80 ap aligned questions

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chemical equilibrium happens when the rates of forward and reverse reactions are equal so the concentrations of products and reactants are constant equilibrium is a dynamic process the concentrations of reactants and products no longer change with time reactions do not stop

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chemical equilibrium condition in the course of a reversible chemical reaction in which no net change in the amounts of reactants and products occurs a reversible chemical reaction is one in which the products as soon as they are formed react to produce the original reactants

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key concepts and summary a reversible reaction is at equilibrium when the forward and reverse processes occur at equal rates chemical equilibria are dynamic processes characterized by constant amounts of reactant and product species glossary

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in a chemical reaction chemical equilibrium is the state in which both the reactants and products are present in concentrations which have no further tendency to change with time so that there is no observable change in the properties of the system 1

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therald moeller clyde metz in chemistry with inorganic qualitative analysis 1980 16 1 reversibility and equilibrium earlier section 4 6 we defined equilibrium as a state of balance between equal and opposing forces any reversible chemical reaction can achieve equilibrium which for such a reaction is a dynamic state

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quiz unit test about this unit enter the dynamic world of chemical reactions explore how factors like temperature and concentration shift reaction rates and witness the balancing act of reversible reactions

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chemical equilibrium is the state of a system in which the rate of the forward reaction is equal to the rate of the reverse reaction figure 1 equilibrium in reaction h 2 g i 2 g 2hi g figure 2 equilibrium between reactants and products

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chemical equilibrium is the state of a system in which the rate of the forward reaction is equal to the rate of the reverse reaction figure 2 equilibrium in reaction h 2 g i 2 g 2hi g chemical equilibrium can be attained whether the reaction begins with all reactants and no products all products and no reactants or some of both

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the description of equilibrium in this concept refers primarily to equilibrium between reactants and products in a chemical reaction other types of equilibrium include phase equilibrium and solution equilibrium a phase equilibrium occurs when a substance is in equilibrium between two states

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before we define chemical equilibrium let s first review the basics of chemical reactions and equations a chemical reaction is a process in which one or more substances are

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review of chemical equilibrium equilibrium experiments study how the concentration of reaction products change as a function of reactant concentrations and or reaction conditions for a typical bimolecular equilibrium reaction such as a b ab increasing amounts of reactant a might be titrated against a fixed amount of the reactant b and

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chemical equilibrium 14 1 the concept of equilibrium and the equilibrium constant many chemical reactions do not go to completion but instead attain a state of chemical equilibrium chemical equilibrium a state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant

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