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salts such as nah contain the hydride ion h when sodium hydride is added to water it produces hydrogen gas in a highly vigorous reaction write a balanced chemical equation for this reaction and identify the conjugate acid base pairs write the expression for K_a for each reaction

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conceptual questions acids bases and conjugates miscellaneous 1 in the brønsted lowry definition of acids and bases an acid a is a proton donor b is a proton acceptor c forms stable hydrogen bonds d breaks stable hydrogen bonds e corrodes metals 2 in the brønsted lowry definition of acids and bases a base

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acids increase the concentration of hydronium ions in solution by donating hydrogen ions to water molecules 2 points true the concentration of hydronium ions is greater

than 1×10^{-7} for acidic solutions
2 points
true what is the pH of a solution with a 4.60×10^{-4} M hydroxide ion concentration
4 points

practice packet unit 12 acids and bases mr palermo s Dec 03 2023

lesson 1 arrhenius acids and bases
1 use table K and table L to help you identify the rules for determining whether a substance is an acid a base or a salt based on the formula
underline all the acids circle bases and box in salts purple leave the covalent substances alone

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acids bases calculations practice worksheet directions solve the following pH calculations write the formula plug numbers into formula give answer with correct units and significant figures
1 if the pH of a solution is 10.3 what is the H⁺ concentration
2 if the H⁺ is 2.1×10^{-12} M HClO₄ what is the pH

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write the balanced chemical equation for the neutralization of phosphoric acid and magnesium hydroxide then state whether the resulting solution will be slightly acidic slightly basic or neutral answer

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as quickly as possible identify the compounds below by type of acid or base and the symbol to be used in an acid or base calculation answer key below possible types of acid or base answers strong acid weak acid strong base weak base lewis acid neither

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we will calculate the ph of the solutions using the following 3 steps for each problem step 1 what is left in solution step 2 what are the equilibrium concentrations of the species in solution step 3 what is the ph of the solution 1 0.1 M HCl
 $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{Cl}^- + \text{H}_3\text{O}^+$

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acid base calculations answers 1 in each of the following solutions determine 3 a 0.030 mol dm⁻³ HCl HCl is a strong acid
 $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{Cl}^- + \text{H}_3\text{O}^+$
b 0.1 M HBr HBr is a strong acid
 $\text{HBr} + \text{H}_2\text{O} \rightleftharpoons \text{Br}^- + \text{H}_3\text{O}^+$
c 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
d 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
e 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
f 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
g 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
h 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
i 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
j 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
k 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
l 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
m 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
n 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
o 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
p 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
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s 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
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u 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
v 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
w 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
x 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
y 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$
z 0.1 M H_2SO_4 completely breaks down therefore $\text{H}_2\text{SO}_4 \rightarrow 2\text{H}^+ + \text{SO}_4^{2-}$

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science chemistry questions and answers date class period homework unit acids bases and solutions acid and base calculations wksh 4.1 calculate the ph of the solutions given

the following h and then identify the solution as acidic basic or neutral

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calculations when a solution is diluted the number of moles of the original solution stays the same even though the volume of the solution changes the following equation is used

$$c_1 V_1 = c_2 V_2$$

c_1 concentration of solution 1 $mol\ dm^{-3}$ c_2 concentration of solution 2 $mol\ dm^{-3}$ V_1 volume of solution 1 dm^3 V_2 volume of solution 2 dm^3

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trigonometry is a branch of mathematics concerned with relationships between angles and side lengths of triangles in particular the trigonometric functions relate the angles of a right triangle with ratios of its side lengths the field emerged in the hellenistic world during the 3rd century bc from applications of geometry to astronomical

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