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assume this is complete combustion 1 $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ 2 $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
3 1 mole C_3H_8 3 moles of CO_2 4 3 moles 2 13 moles 6 39 moles of CO_2
silver tarnishes when exposed to oxygen and dihydrogen monosulfide
the chemical reaction is as follows

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