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solutions are composed of particles the size of an ion or small molecule 0.1–2.0 nm the examples provided above are all considered solutions air is a solution of small gas molecules simple syrup is a solution of sucrose in water and seawater is a solution of ions and water

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a solution which is often abbreviated as soln is defined as a homogeneous mixture containing one solvent and one or more solutes that do not react with one another

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solutions occur frequently in nature and have also been implemented in many forms of manmade technology we will explore a more thorough treatment of solution properties in the chapter on solutions and colloids but here we will introduce some of the basic properties of solutions

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solutions play a very important role in many biological laboratory and industrial applications of chemistry of particular importance are solutions involving substances dissolved in water or aqueous solutions

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in our understanding of chemistry we need to understand a little bit about solutions in this chapter you will learn about the special characteristics of solutions how solutions are characterized and some of their properties

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solution in chemistry a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility the term solution is commonly applied to the liquid state of matter but solutions of gases and solids are possible

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this yields solution processable ni₃ hitp 2 mxene heterostructures with powder and film conductivity of 4 8 10³ and 5 3 10⁵ s m⁻¹ respectively noticeably the excellent solution processability facilitates the fabrication of large area films with good uniformity

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