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use a solubility curve to determine whether a solution is saturated or unsaturated and to estimate the amount of excess solute that remains undissolved in the solution

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the solubility curve plots the changes of the solubility of a solid at different temperatures in a solvent on a graph the variations in temperature are plotted on the x axis and the solubility is plotted on the y axis

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effects of temperature on solubility how to read and interpret a solubility curve examples and step by step demonstration questions and solutions

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here we look at solubility curves we see what they mean how to read them and how to answer questions using them we begin with an overview of solutes solvents solutions and solubility

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the solubility of a solute in a particular solvent is the maximum concentration that may be achieved under given conditions when the dissolution process is at equilibrium when a solute's concentration is equal to its solubility the solution is said to be saturated with that solute

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the solubility curve tells you exactly how many grams of a solute will dissolve in a given amount of solvent and any temperature you can use this curve to answer almost any question you might

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use the solubility curve to answer these questions determine how much potassium nitrate will dissolve in 20 g of water at 50 °C 200 cm³ of saturated lead(II) nitrate solution was prepared at a temperature of 90 °C

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the information that is described above can be visually represented using a solubility curve which graphs the experimentally determined solubility of a solute as a function of its temperature

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base your answers to questions 67 and 68 on the information below the graph shows the relationship between the solubility of a sequence of primary alcohols in water and the total number of carbon atoms in a molecule of the corresponding alcohol at the same temperature and pressure

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a solubility curve is a pictorial representation of the solubility of a solute in a solvent the curve represents the solubility limits of a given mass of solute in grams at a specific

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this chemistry video tutorial provides a basic introduction into solubility curves it contains plenty of examples and practice problems on calculating the

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the concentration of sugar in the solution at this point is known as its solubility the solubility of a solute in a particular solvent is the maximum concentration that may be achieved under given conditions when the dissolution process is at equilibrium

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a solubility curve is a graph of solubility measured in g 100 g water against temperature in c solubility curves for more than one substance are often drawn on the same graph allowing

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solubility curves study the solubility curves in the figure and then answer the questions that follow 1 what relationship exists between solubility and temperature for most of the substances shown 2 a what is the exception nn t i b what general principle accounts for this ex ception 6 6 3 a approximately how many grams of nano₃

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2 solubility curves use the solubility graph to answer the following a 100 ml of saturated solutions of the following salts are prepared at 20 c determine the amount of salt used for the following 1 sodium chloride nacl 37 g 2 potassium chloride kcl 33 g 3 potassium nitrate kno₃ 33 g 4 sodium nitrate nano₃ 85 g b

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use the solubility cure below to answer the following questions 34 which salt is least soluble at 20 c 35

how many grams of kbr can be dissolved in 100g of water at 60 c

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a solubility curve is a data based graph comparing the amount of solute that will dissolve in a given amount of solvent at various temperatures the most typical solubility curves are graphed based solid and gaseous solutes dissolved in 100 grams of water

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use a solubility curve to determine the solubilities of substances at various temperatures explain the effect of pressure on the solubility of gases previously we looked at the primary characteristics of a solution and how water is able to dissolve solid solutes

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the solubility of a substance refers to the amount of solid that dissolves to form a saturated solution usually the units for solubility are in grams per liter molar solubility refers to the number of moles of the solid that dissolve to form one liter of the saturated solution

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