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calculate the number of moles and the mass of the solute in each of the following solutions a 200 l of 18.5 M H₂SO₄ concentrated sulfuric acid b 100.0 ml of 3.8 × 10⁻⁵ M NaCN the minimum lethal concentration of sodium cyanide in blood serum

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write a balanced molecular equation describing each of the following chemical reactions solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas gaseous butane C₄H₁₀ reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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balance the following chemical equation Mg(OH)₂ + HCl → MgCl₂ + H₂O note all reactants and products require a coefficient of at least one learn for free about math art computer programming economics physics chemistry biology medicine finance history and more

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explain your answer and support it with a specific example why is chemical equilibrium described as a dynamic process describe this process in the context of a saturated solution of NaCl in water

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you might need calculator the molecular weight of sodium chloride NaCl is 58.44 g/mol how many

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problem 1 the half life of ^{71}Zn is 2.4 minutes if one had 100.0 g at the beginning how many grams would be left after 7.2 minutes has elapsed solution 7.2 / 2.4 = 3 half lives 100 / 2 = 50 / 2 = 25 / 2 = 12.5 the amount remaining after 3 half lives 100.0 g \times 0.125 = 12.5 g remaining

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reaction rate and the concentration of reactants have a direct relationship as concentration goes up so does reaction rate this is because a higher concentration causes the reactants to have a higher probability to collide with another reactant particle possibly inducing the chemical reaction

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